

# Displacement of Metals

# Chemical 16 Equations

## General Rule



A metal will be displaced from a salt by a **more reactive** metal.

e.g. 1. Magnesium + Lead Oxide  $\rightarrow$  Magnesium Oxide + Lead

A metal will not be displaced from a salt by a **less reactive** metal.

e.g. 2. Lead + Magnesium Oxide  $\rightarrow$  No Reaction

### Task 1

**In your exercise book**, write word equations for the following reactions. If you have been taught how to, then write a balanced symbol equation under each word equation (assume a valency of 3 for iron, 2 for lead and 1 for copper).

	Reactants
a.	Aluminum and lead chloride
b.	Potassium and calcium chloride
c.	Sodium and iron chloride
d.	Zinc and aluminum chloride
e.	Lead and copper chloride
f.	Iron and copper oxide
g.	Calcium and lead iodide
h.	Tin and zinc sulfate
i.	Calcium and copper sulfate
j.	Sodium and potassium fluoride
k.	Magnesium and lead nitrate
l.	Iron and zinc sulfate
m.	Copper and aluminum carbonate
n.	Potassium and magnesium chloride
o.	Zinc and lead sulfate
p.	Magnesium and calcium iodide
q.	Aluminum and lead bromide
r.	Potassium and sodium carbonate
s.	Calcium and copper nitrate
t.	Zinc and copper sulfate

### Reactivity Series

